

ACTION OF DIAZOANISOLE UPON PHENOLES.

By JAMES H. STEBBINS, JR., F. C. S.

SECOND PAPER.

In my last paper on this subject, I described a compound produced by the action of diazoanisole chloride, upon an alkaline solution of orceine.

At present it is my intention to describe a somewhat similar body resulting from the union of diazoanisole chloride, and resorcine disulphonate of soda.

12.3 g. anisidine were diazotized with 17.9 g. H Cl (1.20 Sp. Gr.) and 6.9 gr. Na NO₂, in 200 c. c. water.

The diazoanisole chloride thus formed, was added, little by little, and under constant stirring to a solution composed of 18.9 grs. resorcine disulpho acid, 16 grs. Na OH in 200 c. c. water, and 20 c. c. Na OH and Aq. (10 per cent.)

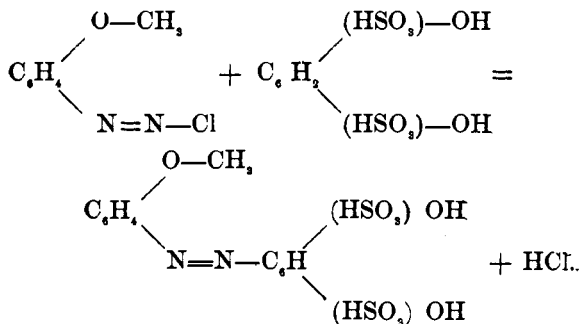
The whole was allowed to rest for a few hours, and then treated with an excess of strong muriatic acid, which precipitated the dye-stuff as a free acid.

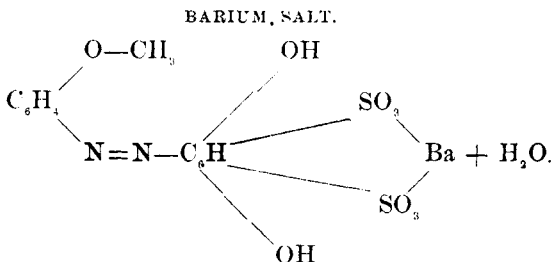
It was then dissolved in hot water and allowed to crystallize.

Thus prepared it crystallizes in little fern shaped yellow plates, which are easily soluble in hot and cold water.

RESORCINE-DISULPHO ACID AZO-ANISOLE.

This substance is probably formed according to the following equation :





This interesting salt, was obtained, by dissolving the free acid in hot water and treating with an excess of Ba Cl₂.

The precipitate formed was collected on a filter, washed and dissolved in boiling water, from which it crystallized on cooling in fine, yellow, star-grouped needles.

In transmitted light this salt has a yellow color, but in reflected light a brick-red shade.

It is sparingly soluble in cold, but readily soluble in boiling water, from which it crystallizes with one molecule water of crystallization.

A barium estimation gave :

Barium.....24.56%

THEORY.

Barium.....24.56%

ON A NEW DEVICE FOR DETERMINATION OF CARBON IN CAST IRON.

BY A. A. BRENEMAN, S. B.

The carbonaceous residue obtained when cast iron is dissolved is commonly transferred with the mass of asbestos serving as a filter to the combustion tube. To avoid accidents incident to this transfer, and to ensure the combustion of refractory portions of graphite, the writer has adopted the modifications of the process described below. A brief note upon the process was read at the meeting of the American Association in 1879, but the success attending its use since then by students under the writer's direction, as well as the introduction of some minor improvements, has led him to believe that a fuller description of the process might be useful to others. The process depends in the main upon the use of combustion tubes of porcelain, and the employment of a small tube of platinum as a